

NOMENCLATURE CHART

OXOACIDS AND THEIR ANIONS

Acids (name of compound)	Anions (form latter part of name of ionic compound)
HNO ₃ nitric acid	NO ₃ ⁻ nitrate
HNO ₂ nitrous acid	NO ₂ ⁻ nitrite
HClO ₄ perchloric acid	ClO ₄ ⁻ perchlorate (<i>Same for Br and I</i>)
HClO ₃ chloric acid	ClO ₃ ⁻ chlorate (<i>Same for Br and I</i>)
HClO ₂ chlorous acid	ClO ₂ ⁻ chlorite (<i>Same for Br and I</i>)
HClO hypochlorous acid	ClO ⁻ hypochlorite (<i>Same for Br and I</i>)
HC ₂ H ₃ O ₂ acetic acid	C ₂ H ₃ O ₂ ⁻ acetate
H ₂ CO ₃ carbonic acid	HCO ₃ ⁻ hydrogen carbonate or bicarbonate CO ₃ ²⁻ carbonate
H ₂ SO ₄ sulfuric acid	HSO ₄ ⁻ hydrogen sulfate or bisulfate SO ₄ ²⁻ sulfate
H ₂ SO ₃ sulfurous acid	HSO ₃ ⁻ hydrogen sulfite or bisulfite SO ₃ ²⁻ sulfite
H ₃ PO ₄ phosphoric acid	H ₂ PO ₄ ⁻ dihydrogen phosphate HPO ₄ ²⁻ hydrogen phosphate PO ₄ ³⁻ phosphate

BINARY & OTHER NON-OXOACIDS AND THEIR ANIONS

HF(aq) hydrofluoric acid	F ⁻ fluoride
HCl(aq) hydrochloric acid	Cl ⁻ chloride
HBr(aq) hydrobromic acid	Br ⁻ bromide
HI(aq) hydriodic acid	I ⁻ iodide
H ₂ S(aq) hydrosulfuric acid	S ²⁻ sulfide
HCN(aq) hydrocyanic acid	CN ⁻ cyanide

MISCELLANEOUS IONS

H ⁻ hydride	NH ₄ ⁺ ammonium
OH ⁻ hydroxide	Ag ⁺ silver
MnO ₄ ⁻ permanganate	Zn ²⁺ zinc
CrO ₄ ²⁻ chromate	Cd ²⁺ cadmium
Cr ₂ O ₇ ²⁻ dichromate	Al ³⁺ aluminum
O ₂ ²⁻ peroxide	
AsO ₄ ³⁻ arsenate	

GREEK PREFIXES (BINARY MOLECULAR CPDS)

1	mono-	6	hexa-
2	di-	7	hepta-
3	tri-	8	octa-
4	tetra-	9	nona-
5	penta-	10	deca-

it's the only non-metal cation we will see!

